

Year 10 Work Plan for the Period Sep 2014 - Dec 2014

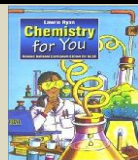
Course: IGCSE Chemistry

Teacher: B.P. Howe

Prescribed Text: Chemistry for You

Support materials: CGP books

Number of periods per week: 5 (3 theory + 2 laboratory)



Week	Date	Theory					Practical No.	Hazards
		Topic	CfY page number	Syllabus section	Manual references	CIE SoW Section		
35	27-Aug	First week, introduction etc						
36	1-Sep	Classifying Materials (Chapter 1) States of matter, invisible particles	6	1	1	10	1.1 - 1.3	C, H, NH ₃
37	8-Sep	Changing state Summary	10-12	1	1	10	1.4 - 1.6	H, Δ
38	15-Sep	Elements, mixtures and compounds (Chapter 2) Atoms, molecules, elements, compounds	14-16	3.2	5	10	2.1, 2.2	SO ₂ , Δ
39	21-Sep	Filtration, evaporation, distillation, chromatography Chemical reactions	18-22	2.2	5	5	2.3, 2.4	F, Δ
40	29-Sep	65 th Year of the Revolution Holiday 						
41	6-Oct	Balancing equations, state symbols Atomic Structure (Chapter 3)	24-28	3.1	11	10	2.5, 2.6	F, Δ
42	13-Oct	Describing atoms, atomic numbers and mass numbers Isotopes, RAM, RFM, Summary	29-36	3.1, 4	11	10	2.7	F, H, Δ
43	20-Oct	The Periodic Table (Chapter 4) groups, metals and non-metals, their properties	42-45	9.1	14	1	4.1	O, Δ
44	27-Oct	Oxides of metals and non-metals, basic and acidic oxides Position in the Periodic Table, Summary	46-49	9.1	14	1	4.2, 4.3	C, FH, SO ₂ , UV, Δ
45	3-Nov	Groups of Metals (Chapter 5) Group I and their reactions, Chemistry at Work	50-53	9.2	20	1	Demo. 5.1, 5.2, 5.3 5.4, 5.5, 5.6	C, HF, FH, Δ C, H ₂
46	10-Nov	Group 2, strontium and barium with water The transition elements	54-57	9.2, 9.3	20	1	Demo 5.7, 5.8, 5.9	C, H ₂ , Δ
47	17-Nov	Chemistry at Work, Summary Groups of Non-metals (Chapter 6)	58-62	9.2, 9.4	24	7 1	6.1 Demo 6.2	C, FH, IR, T, Δ
48	24-Nov	Group 7, reactivity, Chemistry at work Group 8, physical properties, Chemistry at Work	63-68	9.2	24	7	6.3	T, Br ₂
49	1-Dec	Atomic structure and the Periodic Table, Summary The Reactivity Series (Chapter 7) , metals with water	69-80	10.2	28	7 1	Demo 7.1 7.2	C, FH, H ₂ , Δ
50	8-Dec	Metals with acid, displacement reactions, Thermit reaction Displacement from solution, Hydrogen, Summary	81-86	10.2	28	7 1	7.3, 7.4 Demo. 7.5	H, Δ

Notes:

This plan of work is only a guide, but the order will be strictly kept to.

Titles in italics indicate chapter headings from *Chemistry for You*.

It is expected to finish eleven chapters before Easter, if possible.

The pacing of this course will depend on how quickly the students consolidate on their learning.

Extra information is at: www.regoscience.webnode.com

2014

January						February					March						
	1	2	3	4	5	5	6	7	8	9	9	10	11	12	13	14	
Mo		6	13	20	27	3	10	17	24		3	10	17	24	31		
Tu		7	14	21	28	4	11	18	25		4	11	18	25			
We	1	8	15	22	29	5	12	19	26		5	12	19	26			
Th	2	9	16	23	30	6	13	20	27		6	13	20	27			
Fr	3	10	17	24	31	7	14	21	28		7	14	21	28			
Sa	4	11	18	25		1	8	15	22		1	8	15	22	29		
Su	5	12	19	26		2	9	16	23		2	9	16	23	30		
April						May					June						
	14	15	16	17	18	18	19	20	21	22	22	23	24	25	26	27	
Mo		7	14	21	28	5	12	19	26		2	9	16	23	30		
Tu	1	8	15	22	29	6	13	20	27		3	10	17	24			
We	2	9	16	23	30	7	14	21	28		4	11	18	25			
Th	3	10	17	24		1	8	15	22	29	5	12	19	26			
Fr	4	11	18	25		2	9	16	23	30	6	13	20	27			
Sa	5	12	19	26		3	10	17	24	31	7	14	21	28			
Su	6	13	20	27		4	11	18	25		1	8	15	22	29		
July						August					September						
	27	28	29	30	31	31	32	33	34	35	36	37	38	39	40		
Mo		7	14	21	28	4	11	18	25		1	8	15	22	29		
Tu	1	8	15	22	29	5	12	19	26		2	9	16	23	30		
We	2	9	16	23	30	6	13	20	27		3	10	17	24			
Th	3	10	17	24	31	7	14	21	28		4	11	18	25			
Fr	4	11	18	25		1	8	15	22	29	5	12	19	26			
Sa	5	12	19	26		2	9	16	23	30	6	13	20	27			
Su	6	13	20	27		3	10	17	24	31	7	14	21	28			
October						November					December						
	40	41	42	43	44	44	45	46	47	48	49	50	51	52	1		
Mo		6	13	20	27	3	10	17	24		1	8	15	22	29		
Tu		7	14	21	28	4	11	18	25		2	9	16	23	30		
We	1	8	15	22	29	5	12	19	26		3	10	17	24	31		
Th	2	9	16	23	30	6	13	20	27		4	11	18	25			
Fr	3	10	17	24	31	7	14	21	28		5	12	19	26			
Sa	4	11	18	25		1	8	15	22	29	6	13	20	27			
Su	5	12	19	26		2	9	16	23	30	7	14	21	28			

Designed by Anny, annystudio.com

2015

January						February					March						
	1	2	3	4	5	5	6	7	8	9	9	10	11	12	13	14	
Mo		5	12	19	26	2	9	16	23		2	9	16	23	30		
Tu		6	13	20	27	3	10	17	24		3	10	17	24	31		
We		7	14	21	28	4	11	18	25		4	11	18	25			
Th	1	8	15	22	29	5	12	19	26		5	12	19	26			
Fr	2	9	16	23	30	6	13	20	27		6	13	20	27			
Sa	3	10	17	24	31	7	14	21	28		7	14	21	28			
Su	4	11	18	25		1	8	15	22		1	8	15	22	29		
April						May					June						
	14	15	16	17	18	18	19	20	21	22	23	24	25	26	27		
Mo		6	13	20	27	4	11	18	25		1	8	15	22	29		
Tu		7	14	21	28	5	12	19	26		2	9	16	23	30		
We	1	8	15	22	29	6	13	20	27		3	10	17	24			
Th	2	9	16	23	30	7	14	21	28		4	11	18	25			
Fr	3	10	17	24		1	8	15	22	29	5	12	19	26			
Sa	4	11	18	25		2	9	16	23	30	6	13	20	27			
Su	5	12	19	26		3	10	17	24	31	7	14	21	28			
July						August					September						
	27	28	29	30	31	31	32	33	34	35	36	37	38	39	40		
Mo		6	13	20	27	3	10	17	24	31		7	14	21	28		
Tu		7	14	21	28	4	11	18	25		1	8	15	22	29		
We	1	8	15	22	29	5	12	19	26		2	9	16	23	30		
Th	2	9	16	23	30	6	13	20	27		3	10	17	24			
Fr	3	10	17	24	31	7	14	21	28		4	11	18	25			
Sa	4	11	18	25		1	8	15	22	29	5	12	19	26			
Su	5	12	19	26		2	9	16	23	30	6	13	20	27			
October						November					December						
	40	41	42	43	44	44	45	46	47	48	49	50	51	52	53		
Mo		5	12	19	26	2	9	16	23	30		7	14	21	28		
Tu		6	13	20	27	3	10	17	24		1	8	15	22	29		
We		7	14	21	28	4	11	18	25		2	9	16	23	30		
Th	1	8	15	22	29	5	12	19	26		3	10	17	24	31		
Fr	2	9	16	23	30	6	13	20	27		4	11	18	25			
Sa	3	10	17	24	31	7	14	21	28		5	12	19	26			
Su	4	11	18	25		1	8	15	22	29	6	13	20	27			

Designed by Anny, annystudio.com

CIE IGCSE Chemistry Schemes of Work Overview

Scheme of Work	Content
1	Periodic table 1
2	Air and water
3	Acids, bases and salts
4	Organic chemistry 1
5	Chemical analysis techniques
6	Amount of substance
7	Periodic table 2
8	Electrochemistry and redox
9	Organic chemistry 2
10	Chemistry, bonding and structure

Assessment

Paper 1 Multiple choice question paper weighted at 30% of total available marks		45 minutes
and either:		or:
Paper 2 Core theory paper weighted at 50% of total available marks	1 hour 15 minutes	Paper 3 Extended theory paper weighted at 50% of total available marks
and either:		or:
Paper 4 Coursework weighted at 20% of total available marks	Paper 5 Practical test weighted at 20% of total available marks	Paper 6 Alternative to Practical paper weighted at 20% of total available marks

Alterations in the syllabus assessment, content and practical assessment sections for 2011 are indicated by black vertical lines on either side of the text.

Recommended Textbooks

Author	Title	Date	Publisher	ISBN
R. Norris & R. Stanbridge	<i>Chemistry for IGCSE</i>	2009	Nelson Thornes	9781408500187
R. Berry	<i>IGCSE Study Guide for Chemistry</i>	2005	Hodder Murray	0719579023
R. Harwood	<i>Chemistry (Edition 2)*</i>	2003	Cambridge University Press	0521530938

*This book is also available from Cambridge University Press in a Low Priced Edition (ISBN 0 5216 6662 7) from their local distributors in Africa, The Caribbean, Bangladesh, India, Nepal, Pakistan and Sri Lanka.

Reference books

Teachers may also find reference to the following books helpful. They are all suitable for use with this syllabus. Content of the books does not necessarily match the CIE syllabus closely and examples may be British in focus.

Author	Title	Date	Publisher	ISBN
A. Clegg	<i>Chemistry for IGCSE</i>		Heinemann	0435966758
B. Earl & L.D. Wilford	<i>IGCSE Chemistry</i>	2005	Hodder Murray	0719586178
G. Hill	<i>Chemistry Counts</i>		Hodder and Stoughton	0340639342
Lewis & Waller	<i>Thinking Chemistry (GCSE Edition)</i>		Oxford University Press	0199142572

These titles represent some of the texts available at the time of printing this booklet. Teachers are encouraged to choose texts for class use which they feel will be of interest to their students and will support their own teaching style.